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Two equilibria must be considered:  $\text{CuO (s)} + \text{H}_2 \text{(g)} \rightleftharpoons \text{H}_2\text{O (l)} + \text{Cu}$  at equilibrium  $K_1$   $\text{Cu (s)} + \frac{1}{2} \text{O}_2 \text{(g)} \rightleftharpoons \text{CuO}$  at equilibrium  $K_2$  In the final state, only  $\text{O}_2$  is present. Therefore  $K_1 = K_2$ . In an excess of  $\text{O}_2$ , all the 1 copper is oxidized. Therefore  $K_1 = 0$  or  $K_2 = K_1$ .

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of Washington, where he is currently Professor of Chemistry and  
Associate Chair for the Undergraduate Program. Professor Engel  
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## Physical Chemistry Engel Reid Solutions

Thomas Engel has taught chemistry for more than 20 years at the University of Washington, where he is currently Professor of Chemistry and Associate Chair for the Undergraduate Program. Professor Engel received his bachelor's and master's degrees in chemistry from the Johns Hopkins University, and his Ph.D. in chemistry from the University of Chicago.

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